# **Changes in Matter**

Def: A substance changing its form.

There are 4 types of changes: chemical, physical, neutralization and combustion.

#### **Review**

Chemical changes signs	Physical changes signs
colour change	condensation
<ul> <li>heat or light given off</li> </ul>	<ul><li>fusion</li></ul>
<ul> <li>precipitate formed</li> </ul>	<ul><li>sublimation</li></ul>
<ul><li>bubbling</li></ul>	<ul><li>evaporation</li></ul>
<ul><li>mass change</li></ul>	<ul><li>solidification</li></ul>

## **Neutralization**

Def: Occurs when acids and bases are added to each other in the right proportions.

Once neutralization has occurred, acids and bases are no longer present. Instead you have created two new substances.

Reactants Products

Example

NaO/H + ACI 
$$\Rightarrow$$
 H2O + NaCI

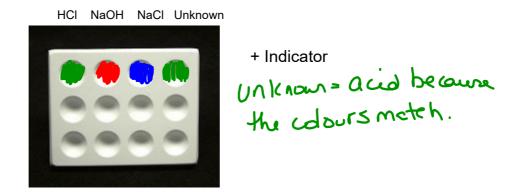
Ca(OHI)<sub>2</sub> +  $32$  HBr  $\Rightarrow$  2 H2O + Ca Br<sub>2</sub>

AI(OH)<sub>3</sub> + 3 HF  $\Rightarrow$  3 H2O + AIF<sub>3</sub>

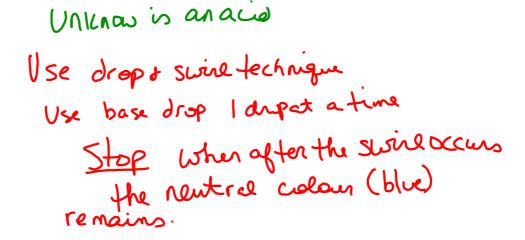
Molecules always produced are? Salt + water

### **Neutralizations labs using indicators**

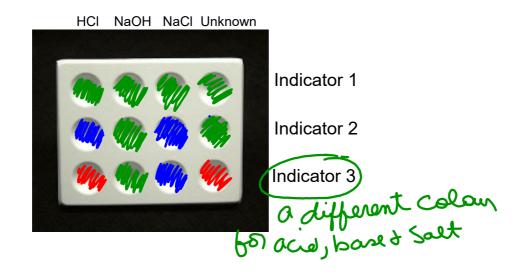
A- How to determine whether an unknown is an acid, base or neutral solution?



B- How to neutralize unknown once its origin is determined?



C- What if you have many indicators to choose from?



## Past exam questions

- 1. In neutralizing sulfuric acid, H<sub>2</sub>SO<sub>4</sub>, with caustic soda, NaOH, sodium sulfate, NA<sub>2</sub>SO<sub>4</sub>, and water are produced. Which equation represents this chemical reaction?
- (A))  $H_2SO_4 + 2 NaOH \rightarrow Na_2SO_4 + 2 H_2O$ 
  - B)  $Na_2SO_4 + 2H_2O \rightarrow H_2SO_4 + 2 NaOH$
  - C)  $H_2SO_4 + NaOH \rightarrow Na_2SO_4 + 2 H_2O$
  - D)  $Na_2SO_4 + H_2O \rightarrow H_2SO_4 + 2 NaOH$